

# **Review of the “Ocean Acidification” *Hypothesis***

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## SUMMARY

Despite a sophisticated effort to promote a new scare about “ocean acidification” that is basically an offshoot of the Global Warming scare, there is no scientific evidence that the new scare has any validity. Our oceans are mildly alkaline or basic with an average pH of about 8 that can vary in time and/or location by about 0.2 up or down. That remarkable stability is maintained by the buffering action of calcium carbonate which is widely abundant. Increasing atmospheric carbon dioxide from whatever source (man-made or natural) will increase the carbonic acid content of the oceans as some of the atmospheric carbon dioxide dissolves in sea water. But the strong buffering action of calcium carbonate will keep the average pH close to 8.0. Hence no significant change in ocean alkalinity is expected. The oceans have not and will not become acidic over large areas as claimed by Al Gore.

The only concern might be the health of some ocean creatures that use calcium carbonate (especially the aragonite crystalline form) for their shells. If there is insufficient calcium otherwise available in a particular area to buffer the additional CO<sub>2</sub> coming into solution, then the sea creatures in that area could lose some shell material to the buffering process. This will not threaten the overall survival of corals because we know they have survived far greater levels of atmospheric carbon dioxide than man can ever create by burning fossil fuels. Corals are very survivable creatures as evidenced by their existence on earth for hundreds of millions of years, far longer than man.

## IMPLICATIONS FOR POLICY MAKERS

An essential conclusion of this paper for policy makers is that they should not reward applicants for grants who use scare tactics in their pursuit of research funds. Doing so only encourages others to use the same approach. Also, it is unlikely that those who exaggerate the value of their research as “absolutely vital” to the survival of the planet will do an honest job with their research.

Ocean research should continue at appropriate levels (not expanded levels), but only with those who do not employ scare tactics. The peer review process needs to be strengthened such that papers for publication are reviewed by competitors of those seeking to publish, not just those who are friends or close associates. Journals (especially Science and Nature) need to enforce rules about data archiving and release, so critics can see if authors are really following accepted standards of academic excellence. Those caught cheating need to be punished. If Science and/or Nature continue favoring one line of reasoning (eg., Global Warming) over all others, their access to federal money should be restricted or terminated.

Climate science needs to be returned to the fundamental dialectic of all science. Theory and observation must be clearly separate pursuits that compliment each other and enforce objectivity. Allowing theory and simulations to become the new standard of rigor is to permit the disintegration of objectivity. Allowing government bureaucrats to decide which theories and simulations are worthy of funding further erodes objectivity. The ultimate arbiter in all questions of science is real evidence obtained by truly independent researchers. The standard of objectivity in all human pursuits is independence from conflicts of interest, especially financial conflicts of interest.

## BACKGROUND

Over the last twenty plus years, we have seen a proliferation of scientific scams based on Al Gore's theory of Anthropogenic Global Warming caused by man-made carbon dioxide. All have similar characteristics, including a thread of truth woven into a narrative about an apocalypse that sounds both convincing and frightening to those *without* scientific training.

To many of those *with* scientific training it has become a dream come true with almost unlimited government funding available. All these scientists need do is profess a belief in Global Warming and convince a government agency that their studies address “the problem.” Of course their studies never address crucial questions such as the role man is actually playing in the changes they observe. Even if they observe changes that clearly indicate cooling rather than warming despite an increase in atmospheric carbon dioxide, they go on as though everything is fine with the underlying hypothesis. Amongst those who owe their jobs to AGW, there is no doubt that Al Gore is correct.

To focus attention on their area of expertise and build public support for their efforts, too many scientists have taken to issuing press releases that highlight their work and point to some “grave concern” about a rapidly approaching “tipping point.” Of course, we are not quite there yet and just might be able to avoid the apocalypse if the government significantly increases their funding. The grave concern could be anything from the melting of the Arctic Tundra, the melting of the Arctic Pack Ice, the melting of the Greenland Ice Cap, warming in West Antarctica, and rising sea levels to a particular species migrating or under stress and the oceans acidifying to the point where much of the life within will be destroyed. The point they wish to drive home is that we just might be able to save the planet if we support their research efforts. The problem is never so dire that all hope is lost or so inconsequential that it can be ignored. Just support more money for them, and the planet will be saved!

Another advantage to having a long list of concerns is purely practical. When the climate stubbornly refuses to warm as the UN IPCC maintains it must, Alarmists can easily shift attention to something else until the climate naturally cycles back to a warm phase.

Just a little knowledge of the natural phenomena under consideration is necessary to spot the essential hoax: on a fluid planet with vast oceans and atmosphere, many physical and biological processes are constantly changing for completely natural reasons. The only thing that would be truly unnatural would be for everything to remain constant. Yet those without scientific training are easily sold on the idea that change is unnatural and must be avoided at all costs. That's rubbish and dangerous.

When finally cornered, many proponents of the apocalypse fall back on the “zero risk” argument. They argue that we should follow their recommendations for reducing per capita carbon emissions to levels not seen since the 19<sup>th</sup> century to eliminate ALL RISK. That betrays a lack of understanding of risk, which is never zero. If we were to return to a primitive 19<sup>th</sup> century lifestyle without the energy that drives our modern economy today, our life expectancy could drop sharply to what it was then. The world as a whole would be far worse off, because a modern mechanized society is necessary to support the world's present population. A return to the past, even if it were possible, is not progress, it is disaster.

## THE ACID OCEAN HYPOTHESIS according to Gore

Here is the hypothesis from Al Gore:

*“Carbon dioxide pollution is changing the very chemistry of our oceans. Ocean acidification is already underway and is accelerating. A recent paper published in the journal Science described how the seawater off the coast of Northern California has become so acidic from CO<sub>2</sub> that it is now corrosive. To give some sense of perspective, for the last 44 million years, the average pH of the water has been 8.2. The scientists at Scripps measured levels off the north coast of California and Oregon at a pH of 7.75. Coral polyps that make reefs and everything that makes a shell are now beginning to suffer from a kind of osteoporosis because of the 25 million tons of CO<sub>2</sub> absorbed by the oceans every 24 hours.”*

This is mostly nonsense. First of all, the oceans are mildly basic or alkaline not acidic, with an average pH of about 8. Figure 1 shows the pH scale with the pH of common solutions indicated. Freshly distilled water is neutral at 7. Rain water is much more acidic at 5.6 because of dissolved CO<sub>2</sub>, beer still more acidic at 4.3, and carbonated beverages still more acidic because they contain phosphoric acid in addition to their carbonation. Soda can be nearly as acidic as lemon juice.

The oceans are almost as basic as a solution of baking soda (sodium bicarbonate) but far from as corrosive as bleach or lye. If the ocean pH is shifting at all, it is shifting ever so slightly toward neutral. There is nothing in the scientific literature to suggest in any way under any circumstances that the oceans will become acidic. That's just Gore's imagination running wild. Scientists do talk about decreases in pH moving the solution toward acidic or increases moving it toward basic. In other words, diluting a solution of sodium hydroxide (lye) could be technically described as acidifying it, but it will never become acidic! Nor should one give the wrong impression by describing conditions at, for instance, Vostok in Antarctic as “warming” when the temperature rises from -100 F to -95 F. It never gets warm there, just less cold.

Figure 2 shows a set of *calculated* values of pH for most of the world's ocean surfaces from Reference 1. As is typical of such charts, the distinction between actual measurements and imaginative simulation is blurred. We assume, for our purposes here, that this comes close to reality, at one point in time.

Note that the lowest pH shown is about 7.9 and the highest about 8.2. There is no “acid” off the Northern California coast and even where the pH of the ocean surface is lowest, it is still solidly basic. In areas where large quantities of fresh water pour into the oceans (such as at the mouth of the Amazon river), it may be possible to observe water that is slightly acidic, because fresh water is generally a little acidic (pH less than 7). Figure 1 suggests that river water has a pH of about 6.

There are indications that corals in some areas suffer from the effects of man – but for much more mundane reasons like mechanical damage, overfishing, water pollution, or sedimentation. (Reference 2).

Humans release about 23 million metric tons of carbon into the atmosphere every day as carbon dioxide. Measurements of atmospheric carbon dioxide show that it is only increasing by about 11 million tons per day. If we assume that all of this increase is human caused (in line with Gore), then 12 million tons must be going elsewhere. There are two general possibilities for “elsewhere:” the

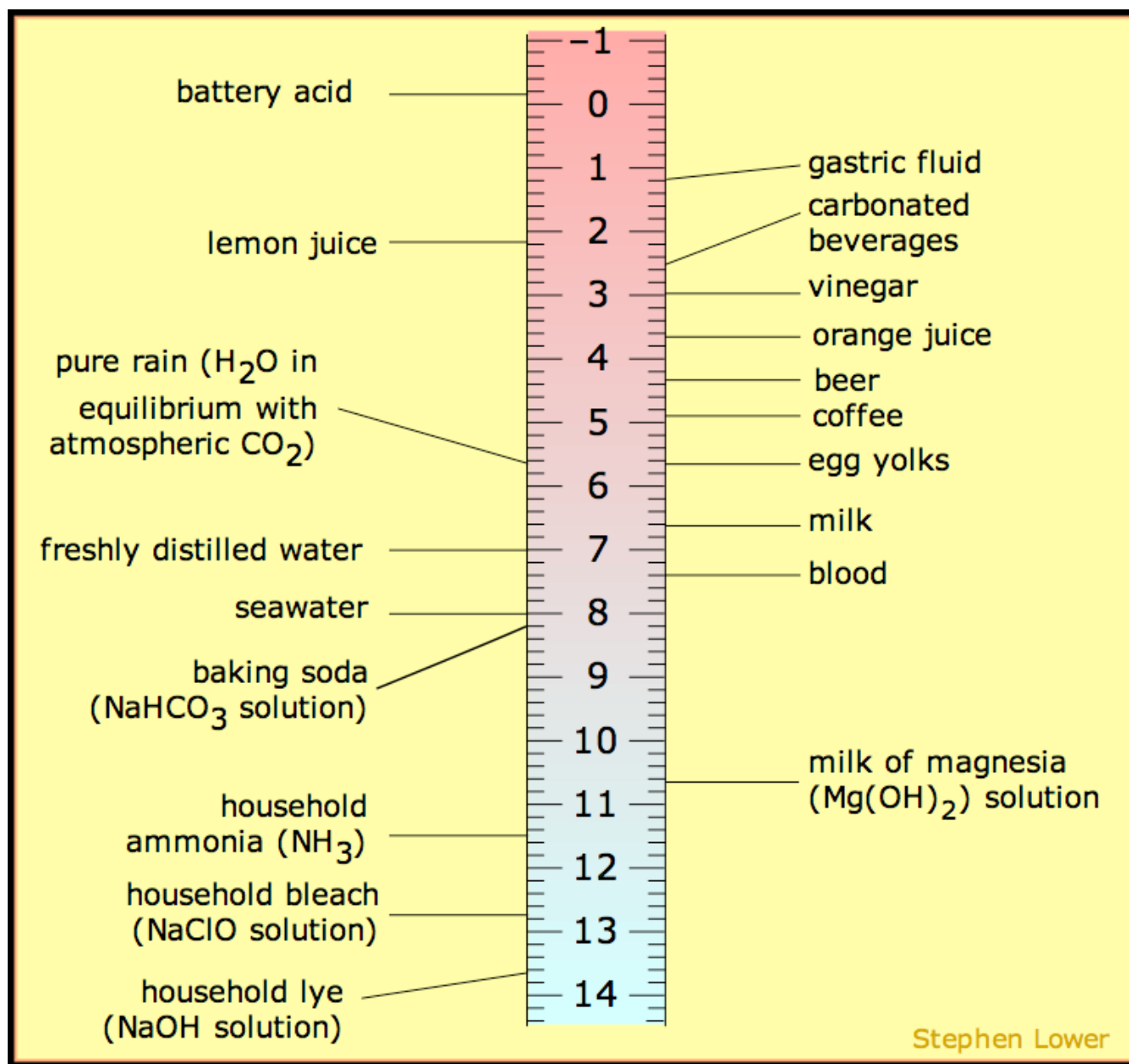


Figure 1. The pH scale showing values for typical aqueous solutions. (from Wikipedia)

biosphere and the oceans. From seasonal variations in atmospheric carbon dioxide, it is clear that the biosphere is capable of absorbing all of this. But for our purposes here, let us assume that roughly half disappears into the biosphere and half into the oceans. Some shows up as increased crop production, for instance, better yields per acre for corn. That leaves about 6 million tons going into the oceans where it likely ends up eventually sequestered in the deep ocean. Six million tons per day may seem like a lot, but is tiny compared to the amounts naturally in play. We humans exhale about 2 million tons of carbon per day just by breathing! If we inflate these numbers by a little more than a factor of three by counting the atmospheric oxygen that goes into making the CO<sub>2</sub>, as Gore does, we still get less than he claims is going into the oceans. In other words, *Gore's numbers are exaggerated.*

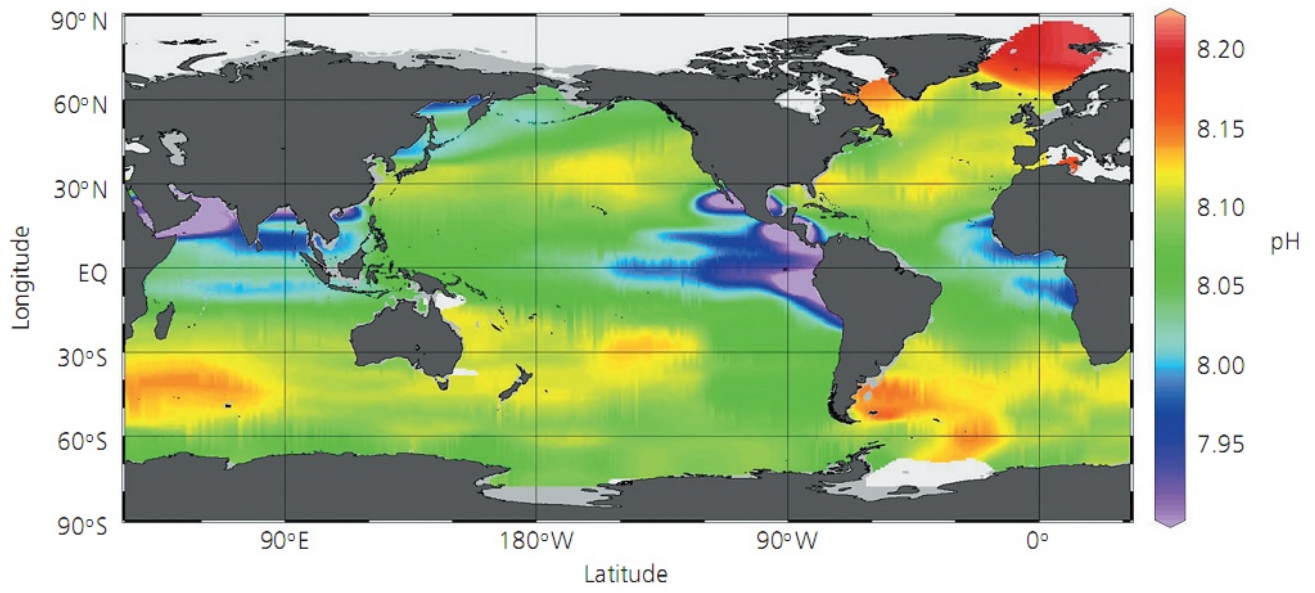


Figure 2. Calculated ocean pH (presumably based on some measurements). (Reference 1)

## THE OCEAN ACIDIFICATION HYPOTHESIS according to the Royal Society

The British Royal Society presents the same hypothesis as Gore but with much more sophistication that makes it seem more plausible:

*“The oceans are absorbing carbon dioxide (CO<sub>2</sub>) from the atmosphere and this is causing chemical changes by making them more acidic (that is, decreasing the pH of the oceans). In the past 200 years the oceans have absorbed approximately half of the CO<sub>2</sub> produced by fossil fuel burning and cement production. Calculations based on measurements of the surface oceans and our knowledge of ocean chemistry indicate that this uptake of CO<sub>2</sub> has led to a reduction of the pH of surface seawater of 0.1 units, equivalent to a 30% increase in the concentration of hydrogen ions.*

*If global emissions of CO<sub>2</sub> from human activities continue to rise on current trends then the average pH of the oceans could fall by 0.5 units (equivalent to a three fold increase in the concentration of hydrogen ions) by the year 2100. This pH is probably lower than has been experienced for hundreds of millennia and, critically, this rate of change is probably one hundred times greater than at any time over this period. The scale of the changes may vary regionally, which will affect the magnitude of the biological effects.*

*Ocean acidification is essentially irreversible during our lifetimes. It will take tens of thousands of years for ocean chemistry to return to a condition similar to that occurring at pre-industrial times (about 200 years ago). Our ability to reduce ocean acidification through artificial methods such as the addition of chemicals is unproven. These techniques will at best be effective only at a very local scale, and could also cause damage to the marine environment. **Reducing CO<sub>2</sub> emissions to the atmosphere appears to be the only practical way to minimise the risk of large-scale and long-term changes to the oceans.***

***All the evidence collected and modelled to date indicates that acidification of the oceans, and the changes in ocean chemistry that accompany it, are being caused by emissions of CO<sub>2</sub> into the atmosphere from human activities. The magnitude of ocean acidification can be predicted with a high level of confidence. The impacts of ocean acidification on marine organisms and their ecosystems are much less certain but it is likely that, because of their particular physiological attributes, some organisms will be more affected than others.....***

***Research into the impacts of high concentrations of CO<sub>2</sub> in the oceans is in its infancy and needs to be developed rapidly. We recommend that a major, internationally coordinated effort be launched to include global monitoring, experimental, mesocosm and field studies. Models that include the effects of pH at the scale of the organism and the ecosystem are also necessary. The impacts of ocean acidification are additional to, and may exacerbate, the effects of climate change. For this reason, the necessary funding should be additional and must not be diverted from research into climate change.”***

The last paragraph in this quotation was set in bold type by the Royal Society and needs to be read with the impact they intended. It is both an admission that very little is really known about the effects of atmospheric CO<sub>2</sub> on the oceans and a plea that governments should set aside huge new sums of money

to fund research in this area. They want the funding to be in addition to the vast funding of Global Warming so as not to be placed in competition with their friends on the Global Warming side. The last thing they want is for Global Warming scientists to argue against their proposed ocean efforts to protect their own research money.

Since the Royal Society paper is a proposal for huge sums of money to conduct new research, since the basic tactic of presenting an apocalyptic view is completely self-serving, and since they never clearly distinguish between what is known based on hard evidence and what is speculative theory, it is hard to take this paper too seriously. Yet this paper is also a semi-competent review article or what passes for one in today's advocacy science. It is one of the primary sources that is quoted today. So let us look at what they say and more importantly what they leave out or ignore that someone not seeking money might argue.

We all agree that atmospheric carbon dioxide combines with water to form a mild acid, carbonic acid  $\text{H}_2\text{CO}_3$ . This makes everything from rainwater to beer slightly acidic. But the hydration equilibrium constant is quite small, so most of the  $\text{CO}_2$  in water remains as  $\text{CO}_2$ . Carbonic acid is important for everything from cave formation in limestone to the transport of  $\text{CO}_2$  and the buffering of pH in mammalian blood. But  $\text{H}_2\text{CO}_3$  is very unstable and readily reverts to  $\text{CO}_2$  and water. Soft drinks, beer, and bubbly wine can be made with  $\text{H}_2\text{CO}_3$  but will easily give up their carbonation if shaken or warmed.

According to Henry's Law the concentration of  $\text{CO}_2$  in the air above an aqueous solution of  $\text{CO}_2$  should maintain a proportionality. In other words, a doubling of atmospheric  $\text{CO}_2$  should result in a doubling of  $\text{CO}_2$  in the aqueous solution at a constant temperature in the absence of buffering. But if the temperature of the aqueous solution goes up, the solution can hold less gas and will out-gas some of its stored  $\text{CO}_2$  to the air above. This is the likely explanation for the observed lag between atmospheric temperature and atmospheric  $\text{CO}_2$  concentration in the ice core data that shows climate history going back almost a million years. The air warms followed by the ocean warming and giving off  $\text{CO}_2$  centuries later.

With  $\text{CO}_2$  and  $\text{H}_2\text{CO}_3$  in sea water, we get bicarbonate ion production as  $\text{H}_2\text{CO}_3$  splits into  $\text{H}^+$  and  $\text{HCO}_3^-$ . But at the same time the carbonate ion present from calcium carbonate ( $\text{CaCO}_3$ ) splitting into  $\text{Ca}^{++}$  and  $\text{CO}_3^{--}$  reacts with  $\text{H}^+$  to produce bicarbonate ion,  $\text{HCO}_3^-$ . This is a classic buffering process. The carbon dioxide entering the water causes a slight decrease in pH followed by a shift back to higher pH by the dissolution of calcium carbonate, and the uptake of hydrogen ions to form bicarbonate ions. The oceans have vast stores of calcium carbonate available for this buffering.

*Hence the ocean pH has to remain very close to 8.0. There is no net acidification even though this scare has been given that name.*

It is remarkable that alarmist scientists continue to use terminology in headlines and summaries that they argue deeper in a paper cannot represent reality.

The concern that they raise is the dissolution of calcium carbonate in the buffering process will harm the ability of sea creatures such as corals to grow or to maintain shells. They admit that the oceans have an abundance of calcium carbonate for the buffering process but assert that ocean mixing is insufficient to make the great stores in sediments available in timely fashion. Hence they argue that sea creatures will suffer as they become the buffering material close to the ocean surface.



To make matters more complicated, calcium carbonate occurs in two crystalline forms in sea creatures: calcite and aragonite. Aragonite is more soluble than calcite; hence sea creatures using this form of calcium carbonate could be more easily harmed in the assumed scenario. Alarmists produce beautiful computer graphics showing *simulated* ocean conditions for aragonite from before the Industrial Revolution (1765) through the future (2100). See Figure 3. Such simulations are likely the source of claims that ocean conditions and pH have changed markedly since the 1700s. There have certainly been no measurements except in very recent times. The pH scale was not even invented until the early 1900s.

Even today there is no reliable, global monitoring of ocean pH, although there have been some global measurements at one point in time as shown in Figure 2. Hence, there is no reliable way to tell whether pH has changed significantly due to any cause, let alone to man.

If measurements had been made in prior centuries, they would likely have shown *different* ocean conditions, much as we know that climate conditions have changed observably over time. But it is ludicrous to attribute these to anthropogenic carbon dioxide in the 1700s, 1800s, or early 1900s because the large increase in man-made carbon dioxide did not occur until after the Second World War. ANY study of varying conditions on earth MUST first eliminate natural causes before attributing variations to man. Failure to do that renders the conclusions invalid.

Anthropogenic carbon dioxide is a minor player in the overall carbon cycle on the earth. Each year man releases about 8.5 Gt of carbon into the atmosphere from the burning of coal, oil, and natural gas as well as the production of cement and breathing. Humans exhale about 0.6 Gt of carbon as CO<sub>2</sub> that originated in plants. (1Gt = 10<sup>15</sup> grams = 10<sup>12</sup> kg = one billion metric tons).

The atmosphere as a whole contains about 780 Gt of carbon; the surface of the oceans contain about 1,000 Gt of carbon; vegetation and soils contain 2,000 Gt of carbon; and the intermediate and deep oceans contain 38,000 Gt of carbon. Each year, the ocean surface and atmosphere exchange an estimated 90 Gt of carbon; vegetation and the atmosphere about 100 Gt of carbon; marine biota and the ocean surface about 50 Gt of carbon; and the ocean surface and the intermediate and deep oceans about 40 Gt of carbon. All of these numbers are greatly larger than the human contributions, indicating that man's effects are small but not completely insignificant. (Reference 3)

Since these natural reservoirs for carbon are highly coupled and constantly interacting in complex ways with far greater exchanges of carbon than man's contribution to the atmosphere, it is not obvious where the small human contribution goes. Measurements of atmospheric carbon dioxide show an increase of only 4 Gt/year of carbon. Hence the remaining human contribution must be absorbed elsewhere. Some (but certainly not all, as the Royal Society maintains) can be going into the oceans. *Much of it is probably going into the biosphere (on land and sea)*. At 388 ppm, the earth's atmosphere has a very low concentration of carbon dioxide that inhibits plant growth. If the concentration falls below 200 ppm, plants cannot grow and will die. They prefer about 1,000 ppm. Recent increases have been enormously beneficial.

From radioisotope studies, we know that an average molecule of carbon dioxide spends about 5 years in the atmosphere. It is frequently estimated that half those molecules leaving the atmosphere go into the ocean surface and half into land vegetation. But this has to vary with ocean surface temperature and growing conditions on land. A warming ocean will naturally out-gas CO<sub>2</sub> (as we see in the ice core data) but should not out-gas as much if man is already augmenting atmospheric CO<sub>2</sub>. In that case, the "excess" man-made CO<sub>2</sub> should move more into vegetation (on land and sea) by benefiting growth

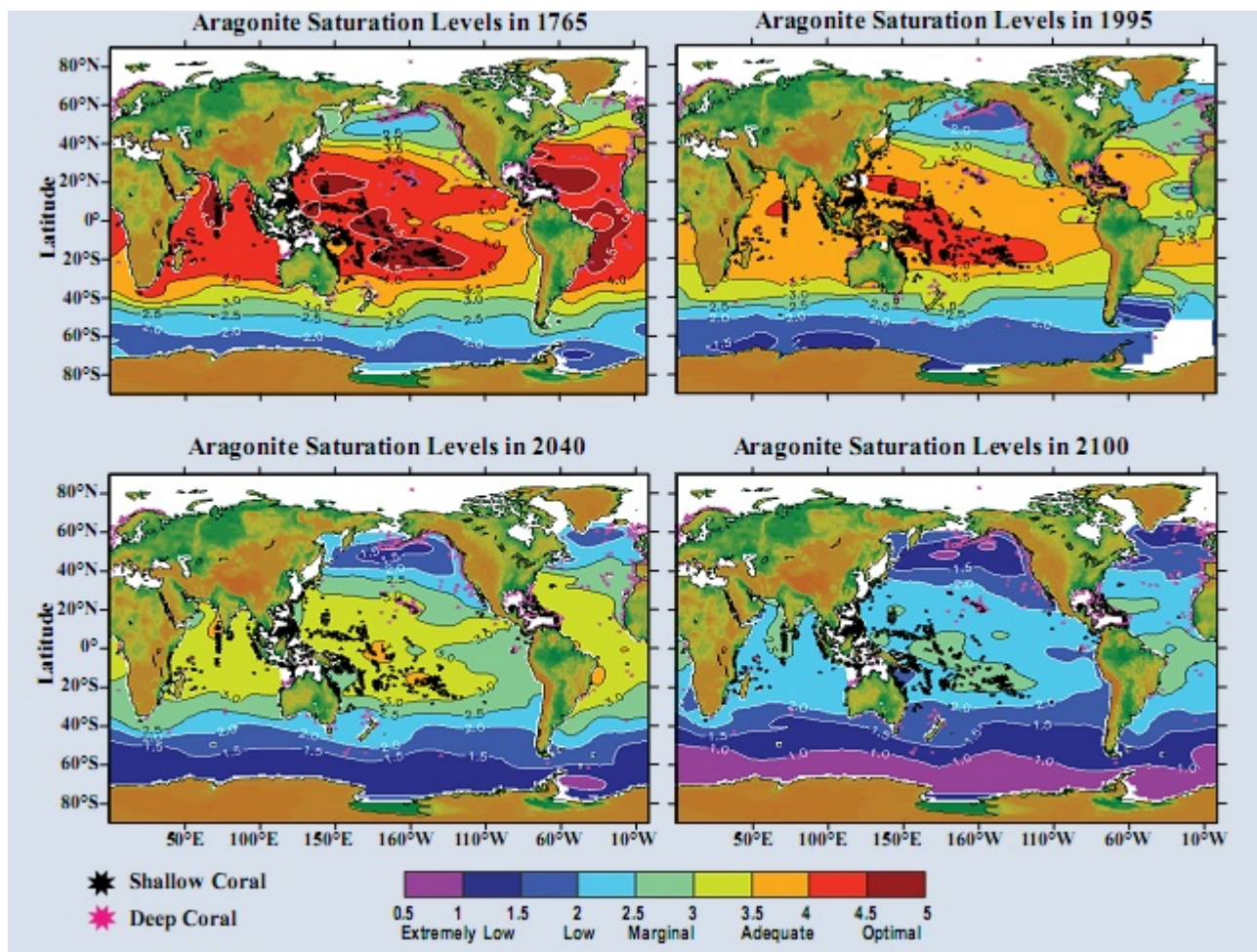


Figure 3 Simulated aragonite saturation states of the surface ocean: past, present, and future. (Feely et al 2006 and Reference 4)

and less into solution in the oceans. Under cooling conditions, the opposite should be true.

The oceans are anything but uniform in temperature and have complex circulation patterns that make the carbon sequestration picture much more complex. But the fact remains that the human part of the carbon cycle is very small and can never become huge because most of the carbon on this planet (that was probably found originally in a primordial atmosphere as CO<sub>2</sub>) is now tied up in non-burnable rocks such as limestone or marble not as coal, oil, or natural gas.

Another perspective for judging the survivability of calcium carbonate creatures like coral under ocean changes that may occur in the future (due to natural or man-made causes) is to look far back in the geological record. Calcite corals are known to have evolved during the Cambrian era, 550 million years ago. At that time there was twenty times as much CO<sub>2</sub> in the atmosphere as there is today. The more delicate aragonite corals evolved 175 million years ago, in the Triassic era, when there was also about twenty times as much CO<sub>2</sub> as there is today. Corals are obviously capable of adapting to huge changes in atmospheric CO<sub>2</sub>, whatever the resulting changes in sea water alkalinity. They have also survived

the large variations in sea level that come with each transition between an ice age and interglacial period. When sea level drops dramatically at the onset of the next ice age, many corals will be left high and dry to die. But those remaining will just set up housekeeping where the new conditions are most appropriate for them. Every creature and plant on earth will do the same thing or perish.

Corals for all their delicate appearance have been extremely resilient over the earth's history, even surviving the geologic event that killed off the dinosaurs. The far more delicate creature, that will have much more trouble surviving natural climate variations certain to occur, is man. Of course, he has a brain but often does not use it.

## **REFERENCES:**

- 1) "Ocean acidification due to increasing carbon dioxide," The Royal Society (policy document), ISBN 0 85403 617 2, June 2005.
- 2) "Humans Are Destroying Earth's Coral Reefs ...," CO2 Science, Volume 12, Number 41: 14 October 2009.
- 3) "Environmental Effects of Increased Atmospheric Carbon Dioxide," Arthur B. Robinson, Noah E. Robinson, and Willie Soon, Oregon Institute of Science and Medicine, 2251 Dick George Road, Cave Junction, Oregon 97523.
- 4) "Impacts of ocean acidification on coral reefs and other marine calcifiers," Report of a workshop sponsored by NSF, NOAA, USGS," JA Kleypas, et. al., June 2006.